

Molecular Geometry

VSEPR Chart



CN	Number of Lone Electron Pairs							
	0	e.g.	1	e.g.	2	e.g.	3	e.g.
2	linear	CO ₂						
3	trigonal planar	BCl ₃ SO ₃	angled	SO ₂ NO ₂ ⁻ O ₃	linear	O ₂		
4	tetrahedral	CH ₄ SO ₄ ²⁻	trigonal pyramidal	NH ₃	angled	H ₂ O	linear	HCl
5	trigonal bipyramidal	PCl ₅	bisphenoidal (seesaw)	SF ₄	T-shaped	ClF ₃	linear	I ₃ ⁻
6	octahedral	SF ₆	square pyramidal	ClF ₅	square planar	ICl ₄ ⁻		
7	pentagonal bipyramidal	IF ₇						